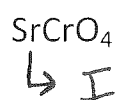


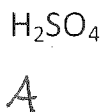
Directions: First identify if the following compounds are ionic (I), molecular (M), or an acid (A). Then write the correct name.



strontium chromate



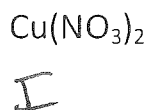
sulfur dibromide



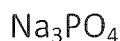
sulfuric acid



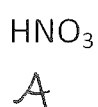
sulfur tetraoxide



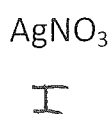
copper (II) nitrate



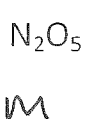
sodium phosphate



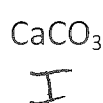
nitric acid



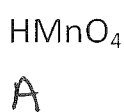
silver nitrate



dinitrogen pentoxide



calcium carbonate



permanganic acid



hydrochloric acid



nitrogen dioxide



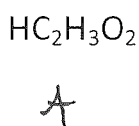
iron (II) oxide



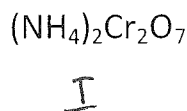
lead (IV) sulfide



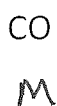
sulfurous acid



acetic acid



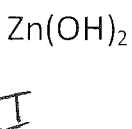
ammonium dichromate



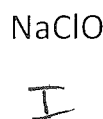
carbon monoxide



potassium chloride



zinc hydroxide



sodium hypochlorite

# ALL MIXED UP

Directions: First identify if the following compounds are ionic (I), molecular (M), or an acid (A). Then write the correct formula.

Calcium oxide

↳ metal - I



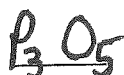
chlorous acid

↳ Acid  $\text{ClO}_2^-$



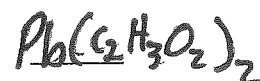
Triphosphorus pentoxide

↳ Non - M



lead (II) acetate

↳ metal - I



Nitric acid

↳ Acid  $\text{NO}_3^-$



aluminum sulfate

↳ metal - I



Sulfur dioxide

↳ Non - M



hydrobromic acid

↳ Acid  $\text{Br}^-$



Ammonium carbonate

↳ Ion - I



silver chromate

↳ metal - I



Potassium permanganate

↳ metal - I



phosphoric acid

↳ Acid  $\text{PO}_4^{3-}$



Magnesium hydroxide

↳ metal - I



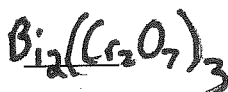
sulfur tetrafluoride

↳ Non - M



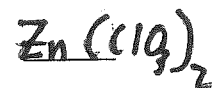
Bismuth (III) dichromate

↳ metal - I



zinc chlorate

↳ metal - I



Sulfuric acid

↳ Acid  $\text{SO}_4^{2-}$



potassium nitrate

↳ metal - I



Sodium bicarbonate

↳ metal - I



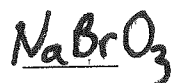
perchloric acid

↳ Acid  $\text{ClO}_4^-$



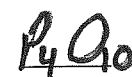
Sodium bromate

↳ metal - I



tetraphosphorus decaoxide

↳ nonmetal - M



$\text{ClO}_3^-$  chlorate